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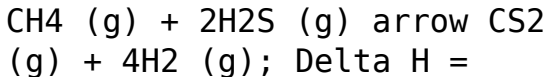
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$10^{-3}$  and  $10^3$ . Chemical  
equilibrium (definition) The  
state reached when the



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concentration of reactants and products remain constant over time. What is the general, reversible reaction for equilibrium? (With letters)  $aA + bB \rightleftharpoons cC + dD$ . Equilibrium equation (general with letters)  $K_c =$

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$[C]^c[D]^d/[A]^a[B]^b$ .

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26 D5 Chemical equilibrium  
is said to be dynamic

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because A. the reaction proceeds quickly. B. the mass of the reactants is decreasing. C. the macroscopic properties are constant. D. both forward and reverse reactions are occurring. 27 D5 Equilibrium

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is a dynamic process because  
the A. macroscopic  
properties are not changing.

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questions 2.5 and 2.6 ||  
\*\*See your instructor to  
find out what questions you  
should complete for review.  
2.3 (a) State in which  
direction the equilibrium  
will shift for each of the

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following, when there are

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This model should show the three states that occur in the chemical equilibrium process. As students work,

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Answers may vary, but should include changes in concentration, volume, and temperature. Some students may list pressure in place of volume. 3. left; remove N<sub>2</sub> or H<sub>2</sub>

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**Answers:** A reversible chemical reaction system can proceed in both directions from a set of initial conditions (concentrations) towards a state of chemical equilibrium.

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The given reaction equation is:  $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$   
 $\Delta H = -191 \text{ kJ}$   
 $2 \text{ S O}_2 (\text{ g}) + \text{ O}_2 (\text{ g}) \rightleftharpoons 2 \text{ S O}_3 (\text{ g}) \Delta H = -191 \text{ k J}$ . The equilibrium constant is a ratio of the forward and reverse ...

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~~Answers~~ Consider the following  
equilibrium reaction. 2 ...  
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1) Give a balanced chemical equation which describes this equilibrium. 2) Explain what influence the removal of  $H_2$  from the container



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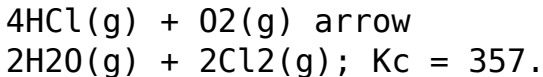
Answers on the  
equilibrium. 3) Explain what  
effect...

~~You are given a container in  
which  $H_2$  (g) ...  
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Solved: Consider the

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following equilibrium:



An equilibrium mixture of the system was found to...

~~Consider the following  
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Answers

Changing the temperature of a reaction at equilibrium alters both the equilibrium constant ( $K$ ) and the equilibrium position. When a reaction is exothermic, which means it releases energy, lowering

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the temperature shifts the  
equilibrium to the  
(14) because the forward  
reaction liberates heat and  
removes the (15).

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