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Chemical Equilibrium

Definition 2. Atomic

Structure Which way

will the Equilibrium

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reaction that feeds the

world - Daniel D. Dulek

The Equilibrium

Constant Chapter 15



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18.1: The Nature of  
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Equilibrium Constant

$K_p$  & Relation in  
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The Laws of

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Chemistry 1. Accuracy  
is the extent to which a  
measurement

approaches the true  
value of a quantity;

precision is the extent to  
which a series of  
measurements of the  
same ...

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Chemical equilibrium is

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a dynamic process consisting of forward and reverse reactions that proceed at equal rates. At equilibrium, the composition of the system no longer changes with time. The composition of an equilibrium mixture is independent of the direction from which equilibrium is approached.

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Chemical reaction is the process by which one or more substances change to produce one or more different substances.

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~~Equilibrium Systems ...~~

Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal and the

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concentrations of the reactants and products remain constant. □

Equilibrium is a dynamic process  $\ominus$  the conversions of reactants to products and products to reactants are still going on, although there is no net change in the number of reactant and product molecules.

~~Chapter 14.~~

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When a chemical  
reaction and its reverse  
occur at the same time  
and rate the reaction is  
said to be in

Equilibrium According  
to Le Chatelier's principle

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will vary but may  
include iron rusting,

milk. . .

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Chemical Equilibrium |

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| N<sub>2</sub> a reversible reaction  
in which there is no

longer any change in the

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concentration of  
reactants and products.  
chemical equilibrium. a  
state of balance in  
which the rate of a  
forward reaction equals  
the rate of the reverse  
reaction and the  
concentrations of  
products and reactants  
remain unchanged.

equilibrium. Holt  
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Cato Maximillian

Galdberg (1836-1902),

and Peter Waage

(1833-1900) Law of

mass action □

relationship between

concentrations of

reactants and products at

equilibrium. If  $aA + bB$

□  $pP + qQ$ . then an

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equilibrium expression  
can be constructed.  $K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$ .

equilibrium expression

depends only on

stoichiometry of

reaction and not

mechanisms.

equilibrium constant:

~~15.S: Chemical~~

~~Equilibrium (Summary)~~

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Our most basic concept

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of equilibrium is based on the observation that change in an isolated system eventually ceases; once change ceases, it never resumes. In this book, we call the idea of a static state of an isolated system the primitive equilibrium. We also observe that change eventually ceases in a closed system that is not



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isolated but whose  
temperature, pressure,  
and volume are kept  
constant.

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