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~~7 Packaging Trends To Watch For In 2019@i-Gong-\u0026-Faichi-with-Dr.-Roger-Jahnke-GMD-Jessica-Koibe-\u0026-Dr.-Roger-Jahnke-on-SBTV-Santa-Barbara PESTO PASTA SALAD RECIPES WITH BACON, MOZZARELLA, AND TOMATOES - Summer Picnic -Fifteen Spatulas STEW challenge results Identifying Anions CSEC Biology SBA Paper 3 Guidelines Become a Qigong and Tai-Chi Teacher EPI Unwrapping the Present Herman Yeung -HKALE Chem - PP 2008/F/Q2b (redox ???? ) Basic Concepts and Innovations in Food Packaging - Smart Packaging Ch 6 - Periodic Table - 6.4 Patterns in the Periodic Table - down the groups (Group VII \u0026 Group 0) Redefining Male Orgasm - Semen Retention Talk LIVE \u0026 Q\u0026A w/ Johnathan White 160112 USMYTH Chemistry SPM: Score A in Chemistry Paper 3 in 15 Minutes??? Jing Kung Chemistry Answer~~

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Question Bank Topic 3 Metals Mastering Chemistry Suggested answers to in-text activities and unit-end exercises Topic 3 Unit 10 Practice P10.1(page 8) a) i) Shiny silvery globules of lead ii) lead(II) oxide + carbon lead + carbon monoxide (or carbon dioxide) b) Carbon monoxide c) No reaction occurs Carbon cannot remove oxygen from aluminium oxide. P10.2(page 12) a) Haematite b) i) The coke ...

#### Master Chem Unit10 (Answers).docx - Question Bank Topic 3 - rrr

Question Bank Topic 3 Metals Mastering Chemistry Heating the oxide of Z gives metal Z. Z is probably a metal low in the reactivity series, such as mercury or silver. 8 A 9 A The more reactive a metal is, the more difficult it is to extract it from its ore. X is a reactive metal as it is difficult to extract from its oxide. Electrolysis of the molten oxide is required.

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New 21st Century Chemistry b) c) d) Discussion (page 177) This discussion aims at raising students' awareness of the principles of green chemistry. See the following list of green chemistry principles the discussion questions help to convey. 1 Safer auxiliary substances - Whenever possible, avoid the use of auxiliary substances such as

#### Suggested answers to in-text activities and exercises

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New 21st Century Chemistry Suggested answers to in-text activities and unit-end exercises Topic 10 Unit 36 In-text activities Checkpoint (page 6) 1 Number of moles of H2 formed = = = 0.00800 mol Average rate of formation of H2 = = 2.00 x 10-4 mol s-1 2

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