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Problems Answer Key A In any

stoichiometry problem, the

first step is always to

calculate the number of

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moles of each reactant present. In this case, we are given the mass of  $\text{K}_2\text{Cr}_2\text{O}_7$  in 1 mL of solution, which we can use to calculate the number of moles of  $\text{K}_2\text{Cr}_2\text{O}_7$  ...

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calculation review. What is

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the mass in grams of  $15 \times 10^{16}$  atoms s. Proudly powered by weeblyweebly. What are the reactants. Chapter 12 stoichiometry. Note that chem 11 problems will be more involved and more rigorous than these.



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Questions 1. Aluminum reacts with oxygen to produce aluminum oxide as follows:

$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  a. If you use 2.3 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  can

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you make? b. If you want 3.9 moles of  $\text{Al}_2\text{O}_3$ , how many moles of  $\text{O}_2$  are needed? 2. In the presence of sulfuric acid, metallic iron forms iron(III) sulfate:  $2\text{Fe} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 3\text{H}_2$

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Terms in this set (...)

Stoichiometry. that portion of chemistry dealing with the numerical relationships in chemical reactions . What is stoichiometry based on?

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the law of conservation of mass. What does stoichiometry involve?  
balancing chemical equations and mole ratios. mole ratio. a conversion factor that relates the number of moles of any two ...

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