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Grams to Grams, Mole Ratio Practice Problems~~

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Answers: Moles and Stoichiometry Practice Problems 1) How many moles of sodium atoms correspond to  $1.56 \times 10^{21}$  atoms of sodium?  
 $1.56 \times 10^{21} \text{ atoms Na} \times 1 \text{ mol Na} = 2.59 \times 10^{-3} \text{ mol Na}$   
 $236.022 \times 10 \text{ atoms Na}$  2) Determine the mass in grams of each of the following: a.  $1.35 \text{ mol of Fe}$   
 $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$  1 mol Fe b.  $24.5 \text{ mol O}$

## Stoichiometry Practice Problems With Answers - 11/2020

Answers to Stoichiometry Practice Free Response Test: 1.  $30.40 \text{ g N}$   
 $1 \text{ mol N} / 14.01 \text{ g N} = 2.17 \text{ moles N}$ .  $69.60 \text{ g O} \times 1 \text{ mol O} / 16.00 \text{ g O} = 4.35 \text{ moles O}$ .  $4.35 / 2.17 = 2.00$   $2.17 / 2.17 = 1.00$  So O is 2X more present than N. Empirical formula must be NO. 2. Mass.

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Remember it is a MC test, use the answers ... e. 12 g 7. How many grams of nitric acid,  $\text{HNO}_3$ , ... Practice Test Ch3 Stoichiometry (page 2 of 2) 19. The mass of element X found in 1.00 mole of each of four different compounds is 28.0 g, 42.0 g, 56.0 g, and 70 g, ...

## Chapter 12 Stoichiometry Test Answer Key

Practice Test Ch 3 Stoichiometry Name Per. Remember it is a MC test,

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use the answers ... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be easiest to balance the equation with mostly whole numbers:  $2\text{NH}_3 + \text{O}_2 \rightarrow 2\text{NO} + 3\text{H}_2\text{O}$  ... 7. c First you must realize this is a limiting reactant problem. Found: 3 Jan 2020 | Rating: 89/100

## Stoichiometry Practice Problems Level 1 Answers

2O Then do some stoichiometry using “ easy math ” 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of  $\text{CO}_2 = 44 \text{ g}$ , and 2 moles of  $\text{H}_2\text{O}$  which is 36 g for a total of 80 g 4. d Balance:  $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$  5. d Balance:  $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

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Chemistry Grade 11 Stoichiometry Worksheets - Teacher ...

Stoichiometry Practice Test - Answer Key. Back to the other Stoichiometry Practice Tests and other General Chemistry Practice Tests. The formation of  $\text{NH}_3$  from  $\text{N}_2$  and  $\text{H}_2$  occurs in 85.0% yield. How many grams of ammonia would be experimentally obtained when 12 g of  $\text{H}_2$  reacts with 20g of  $\text{N}_2$ ?

Stoichiometry Practice Test - Answer Key - SarahChem

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## Chemical Stoichiometry Test Answers

Stoichiometry Practice Test Short Answer: Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation:  $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$  Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide that can be produced from 5.6 mol of Al. 2.

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Determine the amount (in moles) of a product from a given amount of

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one reactant. Determine the amount (in moles) of a product from a given amount of one reactant. If you're seeing this message, it means we're having trouble loading external resources on our website.

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## Chapter 9 Stoichiometry Test Answer Key Modern Chemistry

Stoichiometry Practice Test Short Answer: Aluminum bromide can be prepared by the reaction of aluminum metal with bromine gas shown by the equation:  $2 \text{Al} + 3 \text{Br}_2 \rightarrow 2 \text{AlBr}_3$  Now suppose that 5.6 mol of aluminum reacts with 4.4 mol of bromine. 1. Calculate the mass of aluminum bromide Page 1/8

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WHAT 3 THINGS VALUES IS A MOLE OF A CHEMICAL

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